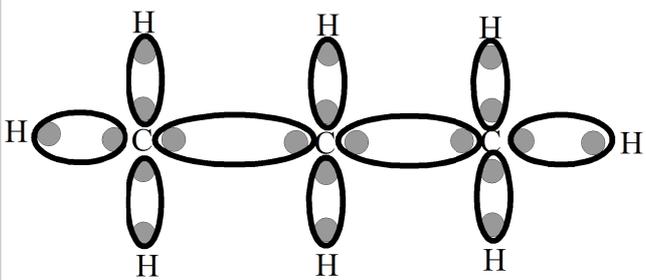
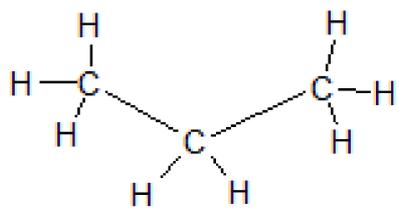


Correction :

Atome	C	H
Structure électronique	$K^2 L^4$	K^1
Nombre d'électrons externes	4 électrons	1 électron
Nombre d'électrons à acquérir (= nombre de liaisons covalentes)	4	1
Représentation de tous les atomes avec leurs électrons externes	 A diagram showing the atomic orbitals of three carbon atoms and eight hydrogen atoms in an ethane molecule. Each carbon atom is represented by a central circle with four lobes (orbitals) extending outwards. Each hydrogen atom is represented by a small circle with one lobe. The orbitals overlap to form covalent bonds, with electron pairs represented by shaded regions at the points of overlap. The text 'Représentation de tous les atomes avec leurs électrons externes' is written to the left of the diagram.	
Formule développée	 A developed structural formula of ethane (C2H6). It shows two carbon atoms (C) connected by a single bond. Each carbon atom is also bonded to three hydrogen atoms (H). The bonds are represented by lines, and the atoms are labeled with their respective symbols. The text 'Formule développée' is written to the left of the diagram.	
Formule semi-développée	$H_3C - CH_2 - CH_3$	